

ICYMI- Chemistry - 11.20.17

1. ACT Practice warm up questions (solubility)
2. Review equations that are applications of the ideal gas equation. (See below.)

$$MM = \frac{gRT}{PV}$$

$$D = \frac{MMP}{RT}$$

3. "Determining Molar Mass Using the Ideal Gas Equation" lab (Part of Unit 10 materials)

Sample data you can use to complete the lab...

DATA TABLE

| | |
|---|-----------------|
| Mass of the lighter before collecting gas | 16.18 g |
| Mass of the lighter after collecting gas | 15.57 g |
| Mass of gas collected | _____ g |
| Volume of gas collected | 250 mL |
| Volume measurement in Liters | _____ L |
| Temperature of the water (and gas) | 20.5 °C |
| Temperature measurement in Kelvins | _____ K |
| Barometric pressure | 30.07 inches Hg |
| Barometric pressure conversion (1 inch = 25.4 mm) | _____ mm Hg |
| Vapor pressure of water at certain temperature * | _____ mm Hg |
| Pressure of the "dry" gas | _____ mm Hg |
| Dry gas pressure measurement in atm | _____ atm |